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Chemical Equilibrium

At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. The equilibrium constant, K , is used to determine the relative concentrations of products and reactants at equilibrium.

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1. $A + B \rightleftharpoons C + D$

(forward reaction) $C +$

$D \rightleftharpoons A + B$ (reverse

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(forward rate = reverse rate) remain constant.
The ratio of the mathematical product $[C]^x \times [D]^y$ to the mathematical product $[A]^n \times [B]^m$ for this reaction has a definite value at a given temperature.

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reaction, indicating how
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sometimes referred to as
the chemical-
equilibrium expression.
The H_2 , I_2 , HI
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Consider the reaction

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18 between H_2 and I_2

vapor in a sealed flask at an elevated

temperature. The rate of reaction can be

followed by observing the rate at which the

violet color of the iodine vapor diminishes, as

$[C]^x[D]^y [A]^n[B]^m$ 556

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Constant Expression.
state reached by a
chemical reaction where
there is no change.... As
equilibrium is

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approached, the rate of
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Explain the nature of the equilibrium constant.

3. Write

chemical equilibrium

expressions and carry

out calculations

involving them. 3. •

Theoretically, every

reaction can proceed in

two directions, forward

and reverse.

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of Reaction. Collision

Theory o Rate = The
speed of any change
that occurs within an

interval of time o KEY

= In chemistry, the rate
of chemical change or
the reaction rate is

usually expressed as the
amount of reactant

changing per unit time o

Collision Theory =

atoms,

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